



## Endothermic and Exothermic Reactions

In a chemical process, the making and breaking of chemical bonds transforms ~~the~~ reactants into products. ~~When two or much more substances counter chemical bond is formed among atoms, creating and forms a chemical compound.~~ Chemical reactions are classified into two types: endothermic and exothermic reactions. Energy activation results in the bonding of two reactants to form a new product.

**Commented [E1]:** The intended meaning of this phrase is difficult to determine, and so it has not been fully revised. Consider rewording the highlighted text to make the meaning of this sentence more apparent.

There are many chemical reactions ~~the that~~ emit energy ~~with in the~~ forms of heat, light, or sound. Such chemical reactions are called exothermic reactions. ~~The is~~ energy that is released comes from the bonds that join several atoms together in the molecules. A common example of ~~an~~ exothermic reactions is ~~the phenomena of~~ combustion. ~~A fully combustible complete~~ ~~In the combustion process, the is a combustible~~ compound reacts with an oxidizing substance, and the compounds of each element in the fuel ~~with the and~~ oxidizing element are emitted. ~~There Most exothermic processes are mostly spontaneous exothermic processes.~~

**Commented [E2]:** I might have unintentionally changed the meaning of this sentence by altering the highlighted text. Please ensure that this phrase still portrays what you intended it to mean.

**Commented [E3]:** I might have unintentionally changed the meaning of this sentence by altering the highlighted text. Please ensure that this phrase still portrays what you intended it to mean.

-On the other hand, many chemical reactions absorb energy ~~in a the~~ heat, light, or sound forms. Such chemical reactions are called endothermic reactions. These reactions cannot progress ~~with without the no~~ addition of heat or ~~supplying~~ energy. The ~~resulting~~ product ~~resulting of from~~ the reaction has ~~lesser low~~ stability because, the higher the energy bond, the ~~less lower the~~ strength ~~its of the~~ molecules ~~possess~~. A common example of ~~an~~ endothermic reactions is ~~the phenomena of~~ photosynthesis. Here, plants use the energy from ~~the~~ sun to convert carbon dioxide and water into glucose and oxygen. Most endothermic reactions are not spontaneous.

To understand the difference between ~~these~~ two reactions types, we need to explore several concepts like the kinetic energy and potential energy behavior in the molecules of the reactants of the chemical reaction.