



## Endothermic and Exothermic Reactions

During a chemical process, ~~the making and breaking of~~ chemical bonds ~~are formed and broken, thereby~~ transform~~ing the~~ reactants into products. ~~When two or much more~~ substances ~~encounter each other, a~~ chemical bond is formed ~~among between these~~ substances' atoms ~~and creating and forms~~ a chemical compound ~~is created~~. ~~There are two~~ broad types of ~~C~~Chemical reactions ~~are classified into two types~~: endothermic and exothermic ~~reactions~~. Energy activation ~~results in the bonding of~~causes two reactants to ~~bond and form~~ a new product.

**Commented [E1]:** I might have unintentionally changed the meaning of this sentence by altering the highlighted text. Please ensure that this phrase still portrays what you intended it to mean.

~~There are many~~Various chemical reactions ~~the~~ emit energy ~~with in the~~ forms of heat, light, or sound. Such chemical reactions are called exothermic reactions. ~~The~~is energy ~~that is~~ released ~~during these reactions~~ comes from the bonds that join several atoms together ~~within the molecules~~. ~~Combustion is a~~ common example of ~~an~~ exothermic reactions, ~~is the~~ phenomena of combustion. ~~A fully combustible~~complete~~In the~~ combustion process, ~~the~~ is a ~~combustible~~ compound reacts with an oxidizing substance, and the compounds of each element in the fuel ~~with the~~and oxidizing element ~~a~~re emitted. ~~There~~ Most exothermic ~~processes~~ are ~~mostly~~ spontaneous ~~exothermic processes~~.

**Commented [E2]:** I might have unintentionally changed the meaning of this sentence by altering the highlighted text. Please ensure that this phrase still portrays what you intended it to mean.

**Commented [E3]:** I might have unintentionally changed the meaning of this sentence by altering the highlighted text. Please ensure that this phrase still portrays what you intended it to mean.

On the other hand, ~~many~~endothermic chemical reactions absorb ~~energy in the~~ heat, light, or sound ~~energy forms~~. Such chemical reactions are called endothermic reactions. These reactions cannot progress ~~with without the~~ no addition of heat or ~~supplying~~ energy. The ~~resulting products of of these~~ reactions ~~has~~ lesser-low stability ~~because, the~~ because of ~~the higher, the~~ energy bonds ~~involved, the less strength~~ its molecules possess. ~~Photosynthesis is a~~ common example of ~~an~~ endothermic reactions ~~is the~~ phenomena of photosynthesis. Here, ~~during which~~ plants use ~~the solar~~ energy ~~from sun~~ to convert carbon dioxide and water into glucose and oxygen. Most endothermic reactions are not spontaneous.

To understand the difference between ~~these~~ two reactions types, we need to explore several concepts, ~~such as~~like the kinetic energy and potential energy behavior ~~in~~ exhibited ~~by~~ the molecules ~~of that comprise~~ the reactants ~~of~~ involved in ~~the~~ chemical reactions.